



MOTHER TERESA WOMEN'S UNIVERSITY
KODAIKANAL – 624 101



DEPARTMENT OF CHEMISTRY

B.Sc., Chemistry

Curriculum Framework, Syllabus and Regulations

**(Based on TANSCHÉ Syllabus under Choice Based
Credit System -CBCS)**



(For the candidates to be admitted from the Academic Year 2023-24)

Mother Teresa Women's University, Kodaikanal
Department of Chemistry
Choice Based Credit System (CBCS)
(2023-2024 onwards)
B. Sc. Chemistry

1. About the Programme

The B.Sc Chemistry Degree Program aims to impart sound knowledge in the fundamental aspects of the important branches of Chemistry. The curriculum is designed to integrate theoretical aspects with experimental/laboratory techniques and analytical thinking which are incorporated in the core and elective courses to equip the learners with the skills required for employability and research. The non-major elective courses, "Clinical chemistry" and "Applied chemistry" provide an overview of the important applications of chemistry to the non-major students. The unique features of the curriculum are ICT based and management oriented skilled based courses, which equip the learners with the essential knowledge of computer applications and managerial skills.

2. Programme Educational Objectives

PEO1	To develop broad knowledge in Chemistry in addition to understanding of key chemical concepts, principles and theories
PEO2	To employ critical thinking and scientific knowledge to design, carry out, record and analyze the results of chemical reactions.
PEO3	To develop students' ability and skill to acquire expertise in solving both theoretical and applied chemistry problems.
PEO4	To provide knowledge and skill to the students' thus enabling them to undertake further studies in Chemistry related areas or multidisciplinary areas that can be helpful for self-employment/entrepreneurship.
PEO5	inculcate the scientific temperament in the students.

3. Eligibility

A candidate who has passed the Higher Secondary Examination with Chemistry, Physics and Mathematics/Zoology as core subjects of Tamil Nadu Higher Secondary Board or an examination of some other board accepted by Mother Teresa Women's University shall be eligible for admission into B.Sc., course in Chemistry.

4. General Guidelines for UG Programme

- i. **Duration:** The programme shall extend through a period of 6 consecutive semesters and the duration of a semester shall normally be 90 days or 450 hours. Examinations shall be conducted at the end of each semester for the respective subjects.
- ii. **Medium of Instruction:** English
- iii. **Evaluation:** Evaluation of the candidates shall be through Internal Assessment and External Examination.

Evaluation Pattern	Theory		Practical	
	Min	Max	Min	Max
Internal	10	25	10	25
External	30	75	30	75

- **Internal (Theory): Test (15) + Assignment (5) + Seminar/Quiz(5) = 25**
 - **External Theory: 75**
- **Question Paper Pattern for External examination for all course papers.**

Max. Marks: 75**Time: 3 Hrs.**

S.No.	Part	Type	Marks
1	A	10*1 Marks=10 Multiple Choice Questions(MCQs): 2 questions from each Unit	10
2	B	5*4=20 Two questions from each Unit with Internal Choice (either / or)	20
3	C	3*15=45 Open Choice: Any three questions out of 5 : one question from each unit	45
Total Marks			75

*** Minimum credits required to pass: 156**

- **Project Report**

A student should carry out and submit the Project Report at the end of the fifth semester. The Project Report shall not exceed 75 typed pages in Times New Roman font with 1.5 line space.

- **Project Evaluation**

There is a Viva Voce Examination for Project Work. The Guide and an External Examiner shall evaluate and conduct the Viva Voce Examination. The Project Work carries 100 marks (Internal: 25 Marks; External (Viva): 75 Marks).

5. Conversion of Marks to Grade Points and Letter Grade

(Performance in a Course/ Paper)

Range of Marks	Grade Points	Letter Grade	Description
90 – 100	9.0 – 10.0	O	Outstanding
80-89	8.0 – 8.9	D+	Excellent
75-79	7.5 – 7.9	D	Distinction
70-74	7.0 – 7.4	A+	Very Good
60-69	6.0 – 6.9	A	Good
50-59	5.0 – 5.9	B	Average
40-49	4.0 – 4.9	C	Satisfactory
00-39	0.0	U	Re-appear
ABSENT	0.0	AAA	ABSENT

6. Attendance

Students must have earned 75% of attendance in each course for appearing for the examination. Students with 71% to 74% of attendance must apply for condonation in the Prescribed Form with prescribed fee. Students with 65% to 70% of attendance must apply for condonation in the Prescribed Form with the prescribed fee along with the Medical Certificate. Students with attendance lesser than 65% are not eligible to appear for the examination and they shall re-do the course with the prior permission of the Head of the Department, Principal and the Registrar of the University.

7. Maternity Leave

The student who avails maternity leave may be considered to appear for the examination with the approval of Staff i/c, Head of the Department, Controller of Examination and the Registrar.

8. Any Other Information

In addition to the above mentioned regulations, any other common regulations pertaining to the UG Programmes are also applicable for this Programme.

Program Outcomes (POs)

On completion of this Programme, the learners will be able to

PO1	To develop broad knowledge in chemistry in addition to understanding of key chemical concepts, principles and theories
PO2	To employ critical thinking and the scientific knowledge to design, carry out, record and analyze the results of chemical reactions.
PO3	To develop students' ability and skill to acquire expertise over solving both theoretical and applied chemistry problems.
PO4	To provide knowledge and skill to the students' thus enabling them to undertake further studies in chemistry in related areas or multidisciplinary areas that can be helpful for self-employment/entrepreneurship.
PO5	To inculcate the scientific temperament in the students.

Program Specific Outcomes (PSOs)

PSO1	Systematic and coherent understanding of the fundamental concepts in Physical chemistry, Organic Chemistry, Inorganic Chemistry, Analytical Chemistry and all other related allied chemistry subjects.
PSO2	Ability to use the evidence based comparative chemistry approach to explain the chemical synthesis and analysis.
PSO3	Ability to demonstrate the experimental techniques and methods of their area of specialization in Chemistry.
PSO4	Developing critical thinking ability by way of solving problems/numerical using Basic chemistry knowledge and concepts
PSO5	Understand good laboratory practices and safety.
PSO6	Inculcating a habit of learning continuously through use of advanced ICT techniques and other available techniques/books/journals for personal academic growth as well as for increasing employability opportunity.

SYLLABUS FRAMEWORK FOR B. Sc CHEMISTRY

(As per TANSCHÉ from 2023-24)

SEMESTER I							
Part	Course Code	Course Title	Credit	No. of Hours	CIA	ESE	Tot. Marks
Part-1	U23CHL11	Language-1 – Tamil	3	6	25	75	100
Part-2	U23CHL12	Language-2 – English	3	6	25	75	100
Part-3	U23CHT11	Core-1: Theory General Chemistry-I	5	5	25	75	100
	U23CHP11	Core-2: Practical Organic Analysis and Organic Estimation	5	5	25	75	100
	U23CHE11	Elective-1: Discipline Specific Elective (Allied) Mathematics I/Zoology	3	4	25	75	100
Part-4	U23CHS1A / U23CHS1B	Skill Enhancement Course SEC – 1: Food Chemistry/ Chemistry in Daily life	2	2	25	75	100
	U23CHF11	Foundation Course	2	2	25	75	100
Total			23	30			

SEMESTER II							
Part	Course Code	List of Courses	Credit	No. of Hours	CIA	ESE	Tot. Marks
Part-1	U23CHL21	Language-1 – Tamil	3	6	25	75	100
Part-2	U23CHL22	Language-2 – English	3	6	25	75	100
Part-3	U23CHT21	Core-3: Theory General Chemistry- II	5	5	25	75	100
	U23CHP22	Core-4: Practical Quantitative Inorganic Estimation and Inorganic Preparation	5	5	25	75	100
	U23CHE21	Elective-2 Discipline Specific Elective (Allied) Mathematics II/Zoology Practical	3	4	25	75	100

Part-4	U23CHS22	Skill Enhancement Course - SEC - 2 (Soft Skills)	2	2	25	75	100
	U23CHS23	Skill Enhancement Course - SEC - 3 Entrepreneurial Skills in Chemistry	2	2	25	75	100
			23	30			

Title of the Course	GENERAL CHEMISTRY-I						
Paper No.	Core 1						
Category	Core	Year	I	Credits	5	Course Code	U23CHT11
		Semester	I				
Instructional hours per week	Lecture	Tutorial	Lab Practice			Total	
	4	1	-			5	
Prerequisites	Higher secondary chemistry						
Objectives of the course	<p>The course aims at giving an overall view of the</p> <ul style="list-style-type: none"> • Various atomic models and atomic structure. • Wave particle duality of matter. • Periodic table, periodicity in properties and its application in explaining the chemical behaviour. • Nature of chemical bonding and fundamental concepts of organic chemistry. 						

Course Outline	<p>UNIT I</p> <p>Atomic structure and Periodic trends History of atom (J.J.Thomson, Rutherford); Moseley's Experiment and Atomic number, Atomic Spectra; Black-Body Radiation and Planck's quantum theory - Bohr's model of atom; The Franck-Hertz Experiment; Interpretation of H-spectrum; Photo electric effect, Compton effect; Dual nature of Matter-De-Broglie wavelength-Davisson and Germer experiment Heisenberg's Uncertainty Principle; Electronic Configuration of Atoms and ions - Hund's rule, Pauli's exclusion principle and Aufbau principle –Numerical problems involving de-Broglie wavelength.</p>
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UNIT II**Introduction to Quantum mechanics**

Classical mechanics, Wave mechanical model of atom, distinction between a Bohr orbit and orbital; Postulates of quantum mechanics; probability interpretation of wave functions, Derivation of Schrodinger wave equation - Probability and electron density-visualizing the orbitals -Probability density and significance of Ψ and Ψ^2 .

Modern Periodic Table

Cause of periodicity; Features of the periodic table; classification of elements - Periodic trends for atomic size- atomic radii, Ionic, crystal and Covalent radii; ionization energy, electron affinity, electronegativity-electronegativity scales, applications of electronegativity.

UNIT III**Structure and bonding-I****Ionic bond**

Lewis dot structure of ionic compounds; properties of ionic compounds; Energy involved in ionic compounds; Born Haber cycle – lattice energies, Madelung constant; Ionpolarization – polarising power and polarizability; Fajans' rules - effects of polarisation on properties of compounds.

Covalent bond

Shapes of orbitals, overlap of orbitals– σ and Π bonds –hybridization of CH_4 , C_2H_4 , C_2H_2 , ; VSEPR theory - shapes of molecules BeCl_2 , H_2O , PCl_3 , NH_3 , CH_4 , PCl_5 , SF_6 .

Partial ionic character of covalent bond-dipole moment, application to molecules of the type $\text{A}_2, \text{AB}, \text{AB}_2, \text{AB}_3, \text{AB}_4$ – percentage ionic character.

UNIT IV**Structure and bonding-II**

VB theory– application to hydrogen molecule; concept of resonance-resonance structures of some inorganic species– CO_2 , NO_2 , CO_3^{2-} , NO_3^- ; limitations of VBT; MO theory-bonding, anti-bonding and non-bonding orbitals, bond order, MO diagrams of H_2 , C_2 , O_2 , O^{2-} , N_2 , CO , NO , HF . Magnetic characteristics, comparison of VB and MO theories.

Coordinate bond: Definition, Adduct formation between BF_3 and NH_3 .

Metallic bond –electron sea model, VB model; Band theory-mechanism of conduction in solids (Brief idea only); conductors, insulator, semiconductor–types, applications of semi-conductors.

Weak Chemical Forces - Vander Waals forces, ion-dipole forces, dipole-dipole interactions, induced dipole interactions, Instantaneous dipole-induced dipole interactions. Hydrogen bonding – Intramolecular and intermolecular hydrogen bonding, special properties of water, ice, stability of DNA (Structure of DNA not needed); Effects of hydrogen bonding on melting and boiling points.

UNIT V**Basic concepts in Organic Chemistry and Electronic effects**

Types of bond cleavage – heterolytic and homolytic- reagents and substrates- types of reagents- electrophiles, nucleophiles, free radicals-reaction intermediates: carbanions, carbo-cations, carbenes and arynes.

Inductive effect - reactivity of alkyl halides, acidity of halo acids, basicity of amines; inductomeric and electromeric effects.

Resonance–resonance energy, conditions for resonance-acidity of phenols, basicity of aromatic amines, stability of carbonium ions, carbanions and free radicals.

Hyperconjugation - stability of alkenes, bond length, orienting effect of methyl group.

Types of organic reactions- addition, substitution, elimination and rearrangements.

Recommended Text

1. Madan, R.D. and Sathya Prakash, *Modern Inorganic Chemistry*, 2nd ed.; S. Chand and Company: New Delhi, 2003.
2. Puri, B. R. and Sharma, L. R. *Principles of Physical Chemistry*, 38th ed.; Vishal Publishing Company: Jalandhar, 2002.
3. Bruce, P.Y. and Prasad K.J.R. *Essential Organic Chemistry*, Pearson Education: New Delhi, 2008.
4. A. Bahl and B.S. Bahl, *Advanced Organic Chemistry, I* Multi color Edition, S. Chand & Company, New Delhi, 2010.
5. Satya Prakash, *Advanced Inorganic*

	Chemistry, R.D. Madan, Vol II, 5 th Edition, S. Chand and Sons, New Delhi, 2012.
Reference Books	<ol style="list-style-type: none"> 1. Maron, S.H. and Prutton C.P. <i>Principles of Physical Chemistry</i>, 4th ed.; The Macmillan Company: New York, 1972. 2. Lee, J.D. <i>Concise Inorganic Chemistry</i>, 4th ed.; ELBS William Heinemann: London, 1991. 3. Gurudeep Raj, <i>Advanced Inorganic Chemistry</i>, 26th ed.; Goel Publishing House: Meerut, 2001. 4. Atkins, P.W. & Paula, J. <i>Physical Chemistry</i>, 10th ed.; Oxford University Press: New York, 2014. 5. Huheey, J.E. <i>Inorganic Chemistry: Principles of Structure and Reactivity</i>, 4th ed.; Addison, Wesley Publishing Company: India, 1993.

Course Learning Outcomes (for Mapping with Pos and PSOs)	
On completion of the course the students should be able to	
CO1	Explain the atomic structure, wave particle duality of matter, periodic properties bonding, and properties of compounds.
CO2	Classify the elements in the periodic table, types of bonds, reactions intermediate electronic effects in organic compounds, types of reagents.
CO3	Apply the theories of atomic structures, bonding, to calculate energy of spectral transitions, Δx , Δp electron gravity, percentage ionic character and bond order.
CO4	Evaluate the relations existing between electronic configuration, bonding, geometry of molecules and reactions; structure activity and electronic effects.
CO5	Construct MO diagrams, predict trends in periodic properties, assess the properties of elements, and explain hybridization in molecules, nature of H – bonding and organic reaction mechanisms.

CO/PSO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3
CO5	3	3	3	3	3
Weightage	15	15	15	15	15
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

Level of Correlation between PSO's and CO's

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M
CO5	S	M	S	S	S	S	S	M	M	S

Title of the Course	ORGANIC ANALYSIS AND ORGANIC ESTIMATION						
Paper No.	Core 2						
Category	Core Practical	Year	I	Credits	5	Course Code	U23CHP11
		Semester	I				
Instructional Hours per week	Lecture	Tutorial		Lab Practice		Total	
	-	-		5		5	
Prerequisites							
Objectives of the course	<p>This course aims at providing knowledge on</p> <ul style="list-style-type: none"> laboratory safety handling glassware analysis of organic compounds organic estimation 						
Course Outline	UNIT I						
	<p>Safety rules, symbols and first aid in chemistry laboratory Basic ideas about Bunsen burner, its operation and parts of the flame. Chemistry laboratory glassware–basic information and uses.</p>						
	<p>Unit II</p> <p>Qualitative Organic Analysis Preliminary examination, detection of special elements-nitrogen, sulphur and halogens Aromatic and aliphatic nature, Test for saturation and unsaturation, identification of functional groups using solubility tests. Confirmation of functional groups</p> <ul style="list-style-type: none"> mono carboxylic acid, dicarboxylic acid monohydric phenol aldehyde, ketone, ester carbohydrate (reducing and non-reducing sugars) primary amine monoamide, diamide. anilide, nitro compound Preparation of derivatives for functional groups 						
<p>UNIT III</p> <p>Organic Estimation</p> <p>a. Estimation of aniline b. Estimation of phenol</p>							

Reference Books	<ol style="list-style-type: none"> 1. Venkateswaran, V.; Veeraswamy, R.; Kulandaivelu, A.R. <i>Basic Principles of Practical Chemistry</i>, 2nded.; Sultan Chand: New Delhi, 2012. 2. Manna, A.K. <i>Practical Organic Chemistry</i>, Books and Allied: India, 2018. 3. Gurtu, J.N.; Kapoor, R. <i>Advanced Experimental Chemistry (Organic)</i>, Sultan Chand: New Delhi, 1987. 4. Furniss, B.S.; Hannaford, A.J.; Smith, P.W.G.; Tatchell, A.R. <i>Vogel's Textbook of Practical Organic Chemistry</i>, 5thed.; Pearson: India, 1989.
Website and e-learning Source	https://www.vlab.co.in/broad-area-chemical-sciences

CO1 Estimate the amount of an organic compound in each solution.

CO2 Identify the presence of special elements and functional group in an unknown organic compound performing systematic analysis.

CO3 Compare mono and dicarboxylic acids, mono and diamides, mono and polyhydric phenols, aldehyde and ketone, reducing & non-reducing sugars and explain the reactions behind it.

CO4 Exhibit a solid derivative with respect to the identified functional group.

CO-PO Mapping (Course Articulation Matrix)

CO/PO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3
Weightage	12	12	12	12	12
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M

Title of the Course	Fundamentals of Chemistry						
Paper No.	Elective 1						
Category	Core	Year	I	Credits	3	Course Code	U23CHE11
		Semester	I				
Instructional hours per week	Lecture	Tutorial	Lab Practice		Total		
	3	1	-		4		

Objectives

1. To understand the handling of chemicals and errors in chemical analysis
2. To get knowledge in chemical bonding and hybridization
3. To acquire knowledge in volumetric analysis
4. To understand the basic concept of chemistry of Thermodynamics and Kinetics

UNIT I

HANDLING OF CHEMICALS AND DATA ANALYSIS

- a) Storage and handling of chemicals: Handling of acids, ethers, toxic and poisonous chemicals. Antidotes, threshold vapour concentration and first-aid procedure.
- b) Errors in chemical analysis: Accuracy, precision. Types of error- absolute and relative errors. Methods of eliminating and minimizing errors.
- c) Separation techniques–Solvent extraction. Principle of adsorption and partition chromatography, column chromatography, thin layer chromatography (TLC), paper chromatography and their applications.

UNIT II

CHEMICAL BONDING

- a) Ionic Bond: Nature of Ionic bond. Structure of NaCl, KCl and CsCl. Factors influencing the formation of ionic bond.
- b) Covalent Bond: Nature of covalent bond. Structure of CH₄, NH₃, H₂O based on hybridization.
- c) Coordinate Bond: Nature of coordinate bond. Coordination complexes. Werner's theory. Geometrical and optical isomerism in square planar and octahedral complexes.
Mention of structure and functions of chlorophyll and hemoglobin
- d) Hydrogen Bond: Theory and importance of hydrogen bonding. Types of hydrogen bonding. Hydrogen bonding in carboxylic acids, alcohol, amides, polyamides, DNA and RNA.

e) van der Waal's forces: Dipole – dipole and dipole - induced dipole interactions.

UNIT III

VOLUMETRIC ANALYSIS

- a) Methods of expressing concentration: normality, molarity, molality, ppm.
- b) Primary and secondary standards: preparation of standard solutions
- c) Principle of volumetric analysis: end point and equivalence points.
- d) Strong and weak acids and bases - Ionic product of water , pH, pKa, pKb. Buffer solutions -pH of buffer solutions. Mention of Henderson equation & its significance.

UNIT IV

KINETICS

- a) Chemical Kinetics: Rate, rate law, order and molecularity. Derivation of rate expressions for I and II order reactions.
- b) Catalysis-Homogeneous and heterogeneous catalysis. Enzyme catalysis, enzymes in biological system and in industry.

UNIT V

THERMODYNAMICS

- a) Introduction: Scope and importance of thermodynamics- system and surrounding-isolated, closed and open systems- state of the system- intensive and extensive variables. Thermodynamic process- reversible and irreversible, isothermal and adiabatic process-
- b) First law of thermodynamics- statement- definition of internal energy (E), enthalpy (H), applications of first law of thermodynamics.

The second law of thermodynamics: Limitations of first law and the need for the second law, different ways of stating II law and its significance, Spontaneous or irreversible process.

The concept of entropy – definition and physical significance of entropy.

Text Books:

1. A.Bahl and B.S. Bahl, Advanced Organic Chemistry, I Multicolor Edition, S.Chand& Company, New Delhi,2010.
- 2.Satya Prakash, Advanced Inorganic Chemistry, R.D. Madan, VolI, 5th Edition, S.Chand and Sons, New Delhi, 2012.
- 3.B.R. Puri, L.R.Sharma and M.S.Pathania, Principles of Physical Chemistry, 46thEdision, Vishal

Publishing Company, New Delhi, 2013.

Reference Book:

1.1.R. Gopalan, S. Sundaram, *Allied Chemistry*, Sultan Chand and Sons, 1995.

CO	Course outcomes	Remarks
CO1	Students can gain the knowledge on the handling of chemicals and errors in chemical analysis.	K2, K3
CO2	Learn Chemical Bonding and Hybridization	K2
CO3	Learn the calculations of preparing standard solutions	K2, K3
CO4	Understand and appreciate the advanced concepts and rate equations in chemical kinetics.	K2
CO5	Calculate change in thermodynamic properties, equilibrium constants, partial molar quantities, chemical potential	K2

K1- Remember **K2-** Understand **K3-** Apply **K4-** Analyze **K5-**Evaluate

Mapping of Cos with POs &PSOs:

PO/PSO	PO1	PO2	PO3	PO4	PO5	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	S	S	S	M	M	M	S	S	S	S
CO2	S	S	S	S	M	M	S	S	S	S
CO3	S	S	S	S	M	M	S	M	S	S
CO4	S	S	S	S	M	M	S	S	S	S
CO5	S	S	S	M	M	M	S	S	S	S

Strongly Correlating(S) - 3 marks

Moderately Correlating (M) - 2 marks

Weakly Correlating (W) - 1 mark

No Correlation (N) - 0 mark

Title of the Course	FOOD CHEMISTRY						
Paper No.	SEC-1						
Category	Core	Year	I	Credits	2	Course Code	U23CHS1A
		Semester	I				
Instructional hours per week	Lecture	Tutorial	Lab Practice		Total		
	2		-		2		
Prerequisites	Higher secondary chemistry						
Objectives of the course	This course aims at giving an overall view of the <ul style="list-style-type: none"> • Types of food • Food adulteration and poisons • Food additives and preservation 						
Course Outline	UNIT I: Food Adulteration Food adulteration-contamination of wheat, rice, milk, butter etc. with claystones, water and toxic chemicals Common adulterants, Ghee adulterants and their detection.						
	UNIT II: Food Poison Food poisons-natural poisons(alkaloids-nephrotoxin)pesticides, (DDT, BHC, Malathion)- Chemical poisons – First aid for poison consumed victims.						
	UNIT III: Food Additives Food additives-artificial sweeteners–Saccharin-Cyclamate and Aspartate Food flavours-some examples–Food colours–Emulsifying agents–preservatives-leavening agents. Baking powder–yeast–tastemakers–MSG-vinegar.						
	UNIT-IV: Beverages Beverages-soft drinks-soda-fruit juices-alcoholic beverages-examples. Carbonation-addiction to alcohol–diseases of liver and social problems.						
	UNIT-V: Edible Oils Fats and oils-Sources of oils-production of refined vegetable oils-preservation. Saturated and unsaturated fats - iodine value - role of MUFA and PUFA in preventing heart diseases.						

	<ol style="list-style-type: none"> 1. Food chemistry, H. K. Chopra, P. S. Panesar, Narosa publishing house, 2010. 2. A text book of pharmaceutical chemistry by Jayashree Ghosh, S Chand publishing, 2012. 3. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006. 4. B. K. Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014. 5. Introduction to forensic chemistry, Kelly M. Elkins, CRC Press Taylor & Francis Group, 2019. 6. Jayashree Ghosh, Fundamental Concepts of Applied Chemistry, S. Chand & Co. Publishers, second edition, 2006.
Reference Books	<ol style="list-style-type: none"> 1. Randolph. Norris Shreve, Chemical Process Industries, McGraw-Hill, Texas, fourth edition, 1977. 2. W. A. Poucher, Joseph A. Brink, Jr. Perfumes, Cosmetics and Soaps, Springer, 2000. 3. A. K. De, Environmental Chemistry, New Age International Public Co., 1990.
Web site and e-learning source	

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M
CO5	S	M	S	S	S	S	S	M	M	S

Course Learning Outcomes (for Mapping with POs and PSOs)
On completion of the course the students should be able to

CO1 Learn about Food adulteration-contamination of Wheat, Rice, Milk, Butter.
CO2 Get an awareness about food poisons like natural poisons (alkaloids-nephrotoxin) pesticides, DDT, BHC, Malathion
CO3 Get an exposure on food additives, artificial sweeteners, Saccharin, Cyclamate and Aspartate in the food industries.
CO4 Acquire knowledge on beverages, soft drinks, soda, fruit juices and alcoholic beverages examples.
CO5 Study about fats and oils-Sources of oils- production of refined vegetable oils-preservation. Saturated and unsaturated fats-MUFA and PUFA

Title of the Course	ROLE OF CHEMISTRY IN DAILY LIFE						
Paper No.	SEC-1						
Category	Core	Year	I	Credits	2	Course Code	U23CHS1B
		Semester	I				
Instructional hours per week	Lecture	Tutorial	Lab Practice		Total		
	2		-		2		
Prerequisites	Higher secondary chemistry						
Objectives of the course	This course aims at providing an overall view of the <ul style="list-style-type: none"> • Importance of Chemistry in everyday life • Chemistry of building materials and food • Chemistry of Drugs and pharmaceuticals 						
Course Outline	UNIT I General survey of chemicals used in everyday life. Air- components and their importance; photosynthetic reaction, green - house effect and the impact on our lifestyle. Water-Sources of water, qualities of potable water, soft and hard water.						
	UNIT II Building materials - cement, ceramics, glass and refractories - definition, composition and application only. Uses of Plastics polythene, PVC, bakelite, polyesters, melamine-formaldehyde resins.						
	UNIT III Food and Nutrition - Carbohydrates, Proteins, Fats -definition and their importance as food constituents– balanced diet–Calories minerals and vitamins (sources and their physiological importance).						
	UNIT IV Chemicals in food production–fertilizers-need, natural sources; urea, NPK fertilizers and super phosphate. Fuel – classification - solid, liquid and gaseous; nuclear fuel examples and uses.						
	UNIT V Pharmaceutical drugs-analgesics and antipyretics-paracetamol and aspirin. pigments and dyes - examples and applications.						

Recommended Text	<ol style="list-style-type: none"> 1. Food chemistry, H. K. Chopra, P. S. Panesar, Narosa publishing house, 2010. 2. A textbook of pharmaceutical chemistry by Jayashree Ghosh, S Chand publishing, 2012. 3. S. Vaithyanathan, Text book of Ancillary Chemistry; Priya Publications, Karur, 2006. 4. B. K. Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014. 5. Introduction to forensic chemistry, Kelly M. Elkins, CRC Press Taylor & Francis Group, 2019. 6. Jayashree Ghosh, Fundamental Concepts of Applied Chemistry, S. Chand & Co. Publishers, second edition, 2006.
Reference Books	<ol style="list-style-type: none"> 1. Randolph. Norris Shreve, Chemical Process Industries, McGraw-Hill, Texas, fourth edition, 1977. 2. W. A. Poucher, Joseph A. Brink, Jr. Perfumes, Cosmetics and Soaps, Springer, 2000. 3. A. K. De, Environmental Chemistry, New Age International Public Co., 1990.
Website and e-learning source	

Course Learning Outcomes (for Mapping with Pos and PSOs) On completion of the course

Course Learning Outcomes (for Mapping with Pos and PSOs) On completion of the course

CO1: Learn about the chemicals used in everyday life as well as air pollution and water pollution.

CO2: Get knowledge on building materials cement, ceramics, glass and plastics, polythene, PVC bakelite, polyesters,

CO3: Acquire information about Food and Nutrition. Carbohydrates, Proteins, Fats Also have an awareness about Cosmetics Toothpastes, face powder, soaps and detergents.

CO4: Discuss about the fertilizers like urea, NPK fertilizers and super phosphate. Fuel classification solid, liquid and gaseous; nuclear fuel-examples and uses

CO5: Have an idea about the pharmaceutical drugs analgesics and antipyretics like paracetamol and aspirin and also about pigments and dyes and its applications

CO1 Learn about the chemicals used in everyday life as well as air pollution and water pollution.
CO2 Get knowledge on building materials cement, ceramics, glass and plastics, polythene, PVC bakelite, polyesters,
CO3 Acquire information about Food and Nutrition. Carbohydrates, Proteins, Fats Also have an awareness about Cosmetics Toothpastes, face powder, soaps and detergents.
CO4 Discuss about the fertilizers like urea, NPK fertilizers and super phosphate. Fuel classification solid, liquid and gaseous; nuclear fuel-examples and uses
CO5 Have an idea about the pharmaceutical drugs analgesics and antipyretics like paracetamol and aspirin and also about pigments and dyes and its applications

CO-PO Mapping (Course Articulation Matrix)

CO/PO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3
CO5	3	3	3	3	3
Weight age	15	15	15	15	15
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

Level of Correlation between PSO's and CO's

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M
CO5	S	M	S	S	S	S	S	M	M	S

Title of the Course	FOUNDATION COURSE						
Category	Foundati on Core	Year Semester	I I	Credits	2	Course Code	U23CHF11
Instructional hours per week	Lecture	Tutorial	Lab Practice		Total		
	2	-	-		2		
Prerequisites	Higher secondary chemistry						
Objectives of the course	<p>This course aims at providing an overall view of</p> <ul style="list-style-type: none"> • Creating interest and confidence in chemistry. • Understanding the fundamental concepts. • Knowing the impact of chemistry in life. 						
Course Outline	UNIT I Remarkable Discoveries and Inventions Major Contributions of Scheele, Lavoisier, Henry Becquerel, Mary Curie, Friedrich Wöhler, Michael Faraday, William Henry Perkin, Ramsay, Sir.C.V.Raman, Haber, Alexander Fleming, Hans Van Pechmann (Only brief idea for each) - Use of MRI scanning, Dialysis in blood purification.						
	UNIT II Laboratory hygiene and safety Reactive inorganic substances and their toxicity (strong acids, bases, halogens, chromates). Hazards due to chemicals, toxic solids, liquids, gases, and other harmful substances - carcinogenic substances. Emergency procedures in chemical splashes to skin and eyes, burns and electric shock. Introduction to lab safety-regulatory requirements-labels, material safety- MSDS. Knowledge of hazard warning information and symbols.						
	UNIT III Principles of volumetric analysis Moles, equivalent weights, Molality, Molarity, Normality, Percentage by Weight and Volume, ppm, - principle of volumetric analysis – primary and secondary standards.						
	UNIT IV Principles of Qualitative Analysis Inorganic qualitative analysis –Common ion effect and solubility product and their application in the precipitation of cations in a mixture.						

	<p>UNIT V Impact of Chemistry in human life Everyday consumer items - Food preservatives, anti-oxidants, (brief idea only) - handmade soaps, shampoo, antiseptics, hair oils, and moisturizer (brief idea only) - Farmyard manure, Compost - Gaseous fuels at home – Glass fibre reinforced plastics and carbon Fibre Reinforced Plastics – examples.</p>
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Recommended Text	<ol style="list-style-type: none"> 1. Elements of Analytical Chemistry by Gopalan Subramanian P.S. Gopalan R., Rangarajan K. Sultan Chand, 2003. 2. Food chemistry, H. K. Chopra, P. S. Panesar, Narosa publishing house, 2010. 3. S. Vaithyanathan, Textbook of Ancillary Chemistry; Priya Publications, Karur, 2006. 4. B.K, Sharma, Industrial Chemistry; GOEL publishing house, Meerut, sixteenth edition, 2014. 5. Introduction to forensic chemistry, Kelly M. Elkins, CRC Press Taylor & Francis Group, 2019.
Reference Books	<ol style="list-style-type: none"> 1. Venkateswaran, V.; Veeraswamy, R.; Kulandaivelu, A.R. Basic Principles of Practical Chemistry, 2nd ed.; Sultan Chand: New Delhi, 2012. 2. Mendham, J.; Denney, R. C.; Barnes, J. D.; Thomas, M.; Sivasankar, B.; Vogel's Textbook of Quantitative Chemical Analysis, 6th ed.; Pearson Education Ltd: New Delhi, 2000.
Web site and e-learning source	<ol style="list-style-type: none"> 1. Timeline of chemistry – Wikipedia. 2. https://www.chemir.com/

Level of Correlation between PSO's and CO's

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M
CO5	S	M	S	S	S	S	S	M	M	S

CO-PO Mapping (Course Articulation Matrix)

CO/PO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3
CO5	3	3	3	3	3
Weightage	15	15	15	15	15
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

Course Learning Out comes (for Mapping with POs and PSOs)

On completion of the course the students should be able to

CO1: Appreciate the evolution of chemistry and the chemists who contributed for chemistry.
CO2: Demonstrate the lab safety-regulatory requirements, procedures in chemical splashes.
CO3: Explain the principles of volumetric analysis.
CO4: Discuss the principles of qualitative analysis.
CO5: Appreciate the impact of chemistry in human life.

SEMESTER II

Title of the Course	GENERAL CHEMISTRY-II						
Paper no.	Core 3						
Category	Core	Year	I	Credits	5	Course Code	U23CHT21
		Semester	II				
Instructional hours per week	Lecture	Tutorial	Lab Practice		Total		
	4	1	-		5		
Prerequisites	General Chemistry I						
Objectives of the course	<p>This course aims at providing an overall view of the</p> <ul style="list-style-type: none"> • chemistry of acids, bases and ionic equilibrium • properties of s and p-block elements • chemistry of hydrocarbons • applications of acids and bases • compounds of main block elements and hydrocarbons 						
Course Outline	<p>UNIT I Acids, bases and Ionic equilibria Concepts of Acids and Bases-Arrhenius concept, Bronsted-Lowry concept, Lewis concept; Relative strengths of acids, bases and dissociation constant; dissociation of poly basic acids, ionic product of water, pH scale, pH of solutions; Degree of dissociation, common ion effect, factors affecting degree of dissociation; acid base indicators, theory of acid base indicators – action of phenolphthalein and methyl orange, titration curves - use of acid base indicators; Buffer solutions – types, mechanism of buffer action in acid and basic buffer, Henderson-Hasselbalch equation.</p> <p>Salt hydrolysis - salts of weak acids and strong bases, weak bases and strong acids, weak acids and weak bases - hydrolysis constant, degree of hydrolysis and relation between hydrolysis constant and degree of hydrolysis. Solubility product - determination and applications. Numerical problems involving the degree of hydrolysis and solubility product.</p>						
	<p>UNIT II Chemistry of s - Block Elements Hydrogen: Position of hydrogen in the periodic table. Alkali metals: Comparative study of the elements with respect to oxides, hydroxides, halides, carbonates and bicarbonates. Diagonal relationship of Li with Mg. Preparation, properties and uses of NaOH, Na₂CO₃, KClO₃- Alkaline earth metals-General group discussion. Chemistry of p- Block Elements (Group 13 & 14) Preparation and structure of diborane and borazine. Chemistry of borax. Extraction of Al and its uses. Alloys of Al. Comparison of carbon with silicon. Carbon-di-sulphide – Preparation, properties, structure and uses. Structure and uses of percarbonates.</p>						

UNIT III**Chemistry of p-Block Elements (Group15-18)**

General characteristics of elements of Group15; chemistry of $\text{H}_2\text{N-NH}_2$, HN_3 and HNO_3 . Chemistry of oxy acids of phosphorous (H_3PO_3 and H_3PO_4).

General properties of elements of group16-Structure and allotropy of elements – Preparation, properties and structure of ozone - Classification of oxides – Oxy acids of sulphur (Caro's and Marshall's acids).

Chemistry of Halogens: General characteristics of halogen with reference to electro-negativity, electron affinity, oxidation states and oxidizing power. Peculiarities of fluorine. Comparative study of halogen acids (HF , HCl , HBr and HI). Uses of HClO_4 . Basic concepts of Inter-halogen compounds- Structure of ICl , ClF_3 , BrF_5 and IF_7 - Basic nature of Iodine.

Noble gases: Position in the periodic table.

Structure of XeF_2 , XeF_4 , XeF_6 and XeOF_4 ; uses of noble gases- clathrate compounds.

UNIT IV**Hydrocarbon Chemistry-I**

Petroproducts: Fractional distillation of petroleum; cracking, isomerisation, alkylation, reforming and uses.

Alkenes-Nomenclature, general methods of preparation – Mechanism of β -elimination reactions – E1 and E2 mechanism - Hofmann and Saytzeff rules. Reactions of alkenes – addition reactions – Markownikoff's rule, Kharasch effect,

oxidation reactions – hydroxylation, oxidative degradation, epoxidation, ozonolysis, polymerization.

Alkadienes

Nomenclature - classification – isolated, conjugated and cumulated dienes– Diels–Alder reactions.

Alkynes

Nomenclature; general methods of preparation, properties and reactions; acidic nature of terminal alkynes and acetylene.

Cycloalkanes: Nomenclature, Relative stability of cycloalkanes, Bayer's strain theory and its limitations.

	<p>UNIT V</p> <p>Hydrocarbon Chemistry-II</p> <p>Benzene: Source, structure of benzene, stability of benzene ring, molecular orbital picture of benzene, aromaticity, Huckel's (4n+2) rule and its applications. General mechanism of aromatic electrophilic substitution-nitration, sulphonation, halogenation, Friedel-Craft's alkylation and acylation. Mono substituted and di substituted benzene- Effect of substituent-orientation and reactivity.</p> <p>Polynuclear Aromatic hydrocarbons: Naphthalene – nomenclature, Haworth synthesis- reactions-preferential substitution at β -position-reduction, oxidation-uses.</p>
Extended Professional Component (is a Part of internal component only, Not to be included in the external examination question paper)	Questions related to the above topics, from various competitive examinations UPSC/JAM/TNPSC others to be solved (To be discussed during the Tutorial hours)
Skills acquired From this course	Knowledge, Problem solving, Analytical ability, Professional Competency, Professional Communication and Transferable skills.
Recommended Text	<ol style="list-style-type: none"> 1. Madan R D, Sathya Prakash,(2003),Modern Inorganic Chemistry, 2nded, S.Chand and Company, New Delhi. 2. Sathya Prakash, Tuli G D,Basu S K and Madan R D, (2003),Advanced Inorganic Chemistry,17thed.,S.Chand and Company, New Delhi. 3. Bahl BS, Arul Bhal,(2003),Advanced Organic Chemistry,3rded., S. Chand and Company, New Delhi. 4. Tewari KS, Mehrothra SN and Vishnoi NK,(1998), Text book of Organic Chemistry ,2nded., Vikas Publishing House, New Delhi. 5. Puri BR, Sharma LR, (2002), Principles of Physical Chemistry,38thed., Vishal Publishing Company, Jalandhar.
Reference Books	<ol style="list-style-type: none"> 1. Maron S Hand Prutton CP, (1972), Principles of Physical Chemistry, 4th ed., The Macmillan Company, New york. 2. Barrow GM,(1992), Physical Chemistry,5thed., Tata McGraw Hill, New Delhi. 3. Lee JD, (1991),Concise Inorganic Chemistry,4thed., ELBS William Heinemann, London. 4. Huheey JE,(1993), Inorganic Chemistry: Principles of Structure and Reactivity, 4thed., Addison Wesley Publishing Company, India. 5. Gurudeep Raj, (2001), Advanced Inorganic Chemistry Vol-I,26thed., Goel Publishing House, Meerut. 6. Agarwal OP,(1995), Reactions and Reagents in Organic Chemistry,8thed., Goel Publishing House, Meerut.

Website and e-learning source	<p>https://onlinecourses.nptel.ac.in/http://cactus.dixie.edu/sblack/chem1010/lecture_notes/4B.htmhttp://www.auburn.edu/~deruija/pdareson.pdfhttps://swayam.gov.in/course/64 -atomic-structure-and-chemical-bonding</p> <p>MOOC components http://nptel.ac.in/courses/104101090/ Lecture 1: Classification of elements and periodic properties http://nptel.ac.in/courses/104101090/</p>
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CO1	Explain the concept of acids, bases and ionic equilibria; periodic properties of s and p-block elements, preparation and properties of aliphatic and aromatic hydrocarbons.
CO2	Discuss the periodic properties of s and p-block elements, reactions of aliphatic and aromatic hydrocarbons and strength of acids.
CO3	Classify hydrocarbons, types of reactions, acids and bases, examine the properties and p-block elements, reaction mechanisms of aliphatic and aromatic hydrocarbons.
CO4	Explain theories of acids, bases and indicators, buffer action and important compounds of s-block elements.
CO5	Assess the application of hard and soft acids indicators, buffers, compounds of s and p-block elements.

Level of Correlation between PSO's and CO's

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M
CO5	S	M	S	S	S	S	S	M	M	S

CO-PO Mapping (Course Articulation Matrix)

CO/PO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3

CO5	3	3	3	3	3
Weight age	15	15	15	15	15
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

Title of the course	QUANTITATIVE INORGANIC ESTIMATION AND INORGANIC PREPARATION						
Paper no.	Core IV						
Category	Core Practicals	Year Semester	I II	Credits	5	Course Code	U23CHP22
Instructional Hours per week	Lecture	Tutorial	Lab Practice		Total		
	-	-	5		5		
Prerequisites	Higher secondary chemistry						
Objectives of the course	<p>This course aims at providing knowledge on</p> <ul style="list-style-type: none"> • Laboratory safety • Handling glasswares • Quantitative estimation • Preparation of inorganic compounds 						
Course outline	<p>UNIT I</p> <p>Common Apparatus Used in Quantitative Estimation (Volumetric) Description and use of burette, pipette, standard flask, measuring cylinder, conical flask, beaker, funnel, dropper, clamp, stand, wash bottle, watch glass, wire gauge and tripod stand.</p> <p>Principle of Quantitative Estimation (Volumetric) Equivalent weight of an acid, base, salt, reducing agent, oxidizing agent; concept of mole, molality, molarity, normality; primary and secondary standards, preparation of standard solutions; theories of acid-base, redox, complexometric, iodimetric and iodometric titrations; indicators—types, theory of acid–base, redox, metal ion and adsorption indicators, choice of indicators.</p>						

	<p>UNIT II</p> <p>Quantitative Estimation (Volumetric) Preparation of standard solution, dilution from stock solution</p> <p>Acidimetry and alkalimetry Titration acids: hydrochloric acid, sulphuric acid Standard solutions prepared: sodium carbonate, sodium bicarbonate, oxalic acid.</p> <p>Permanganometry Estimation of ferrous sulphate and oxalic acid using standard ferrous ammonium sulphate.</p> <p>Dichrometry Estimation of ferric alum using standard dichromate (external indicator) Estimation of ferric alum using standard dichromate (internal indicator)</p>
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	<p>Iodometry Estimation of copper in copper sulphate using standard dichromate</p> <p>Argentimetry (Demonstration only) Estimation of chloride in barium chloride using standard sodium chloride/ Estimation of chloride in sodium chloride (Volhard's method)</p>
	<p>UNIT III Complexometry Estimation of hardness of water using EDTA</p> <p>Preparation of Inorganic compounds Tetra ammine copper(II) sulphate Mohr's Salt</p>
Skills acquired From this course	Knowledge, Problem solving, Analytical ability, Professional Competency, Professional Communication and Transferable skills.
Recommended Text	<p>Reference Books:</p> <ol style="list-style-type: none"> 1. Venkateswaran, V.; Veeraswamy, R.; Kulandivelu, A.R. <i>Basic Principles of Practical Chemistry</i>, 2nded.; Sultan Chand & Sons: New Delhi, 1997. 2. Nad, A.K.; Mahapatra, B.; Ghoshal, A.; <i>An advanced course in Practical Chemistry</i>, 3rded.; New Central Book Agency: Kolkata, 2007.
Reference Books	1. Mendham, J.; Denney, R.C.; Barnes, J.D.; Thomas, M.; Sivasankar, B.; <i>Vogel's Textbook of Quantitative Chemical Analysis</i> , 6 th ed.; Pearson Education Ltd: New Delhi, 2000.
Website and e-learning source	<p>Web References:</p> <ol style="list-style-type: none"> 1) http://www.federica.unina.it/agraria/analytical-chemistry/volumetric-analysis 2) https://chemdictionary.org/titration-indicator/

Course Learning Outcomes (for Mapping with Pos and PSOs)

On successful completion of the course the students should be able to

CO1 Explain the basic principles involve Titrimetric analysis and inorganic preparations.

CO2 Compare the methodologies of different titrimetric analysis

CO3 Calculate the concentrations of unknown solution in different ways and develop the skill estimated the amount of a substance present in a given solution.

CO4 Assess the yield of different inorganic preparations and identify the end point of various titrations.

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M
CO4	S	S	S	S	S	S	S	M	M	M

CO-PO Mapping (Course Articulation Matrix)

CO/PSO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
CO3	3	3	3	3	3
CO4	3	3	3	3	3
Weightage	12	12	12	12	12
Weighted percentage of Course Contribution to Pos	3.0	3.0	3.0	3.0	3.0

Title of the Course	VOLUMETRIC ANALYSIS						
Paper No.	Allied practical						
Category	Skill Enhancement Course	Year	I	Credits	2	Course Code	U23CHE21
		Semester	II				
Instructional Hours per week	Lecture	Tutorial	Lab Practice		Total		
	1	-	1		2		
Prerequisites	Higher Secondary						

This course aims at providing knowledge on

- laboratory safety
- handling glasswares
- Volumetric analysis

Acidimetry and alkalimetry: Titration acids used: hydrochloric acid, sulphuric Standard solutions prepared: sodium carbonate, sodium bi carbonate, oxalic acid.

Oxidation and reduction titration: Oxidising agents: Potassium permanganate (permanganometry) Reducing agents: Ferrous sulphate, ferrous ammonium Sulphate, oxalic acid

Standard solutions prepared: Ferrous Sulphate, ferrous ammonium Sulphate and oxalic acid.

Iodometry titrations: titrations of liberated iodine against sodium thiosulphate using acidified potassium permanganate, potassium dichromate and copper Sulphate solutions.

Standard solutions: potassium dichromate, copper sulphate.

Text Books

1. Sundaram, Krishnan, Raghavan, Practical Chemistry (Part II), S. Viswanathan Co. Pvt., 1996.
2. B.S. Furniss, A.J. Hannaford, P.W. G. Smith, A.R. Tatchell, Vogel's Text Book of Practical Organic Chemistry. 5th Edn., Pearson Education, 2005.

Reference Books

1. Practical Chemistry by A.O. Thomas, Scientific Book Centre, Cannanore, 2003.
2. Basic Principles of Practical Chemistry, V. Venkateswaran, R. Veeraswamy, A. R. Kulandaivelu, Sultan Chand & Sons, New Delhi, 2nd Edn., 2004.

CO	Course outcomes	Remarks
CO1	Learn the concept of Titration methods and various Titrations	K2
CO2	Understand the Acidimetry and alkalimetry titrations	K2
CO3	The preparation of standard solutions and methods of analyze the various salts	K2, K4
CO4	Understand the calculations of molarity, molality and Normality of the solutions	K2

Title of the Course	ENTREPRENEURIAL SKILLS IN CHEMISTRY
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Paper No.	Skill Enhancement Course 3						
Category	Skill Enhancement Course	Year	I	Credits	2	Course Code	U23CHS23
		Semester	II				
Instructional Hours per week	Lecture	Tutorial	Lab Practice		Total		
	1	-	1		2		
Prerequisites	General Chemistry						
Objectives of the course	This course aims at providing training to <ul style="list-style-type: none"> • Develop entrepreneur skills in students. • To provide hands on experience to prepare and develop products. • Develop startups 						
Course Outline	UNIT I MILK AND MILK PRODUCTS Composition of Milk- Flavour and aroma of Milk - Effect of heat on milk ; Pasteurisation ; Homogenisation - Some Milk products : Cheese, Paneer, Cream, Butter, Ice Cream, Milk powder, Curd and butter milk.						
	UNIT II Hands on Experience (Students can choose any four) Detection of adulterants in food items like coffee, tea, pepper, chilli powder, turmeric powder, butter, ghee, milk, honey etc., by simple techniques. Preparation of Jam, squash and Jelly, Gulkand, cottage cheese. Preparation of products like candles, soap, detergents, cleaning powder, shampoos, pain balm, toothpaste/ powder and disinfectants in small scale. Testing of water samples using a testing kit. Dyeing–Cotton fabrics with natural and synthetic dyes Printing–tie and dye, batik.						

Skills acquired From this course	Entrepreneurial skills.
Recommended Text	1. George S & Muralidharan V,(2007) Fibre to Finished Fabric–A Simple Approach, Publication Division, University of Madras, Chennai. 2. Appaswamy G P, A Handbook on Printing and Dyeing of Textiles.
Reference Books	Shyam Jha, Rapid detection of food adulterants and contaminants(Theory and Practice), Elsevier,e-BookISBN9087128004289,1 st Edition,2015
Website and e-learning source	https://www.vlab.co.in/broad-area-chemical-sciences
Course Learning Outcomes (for Mapping with POs and PSOs)	
On completion of the course the students should be able to	
CO1: Identify adulterated food items by doing simple chemical tests.	
CO2: Prepare cleaning products and become entrepreneurs	
CO3: Educate others about adulteration and motivate them to become entrepreneurs.	

	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10
CO1	S	S	S	S	S	S	S	M	S	M
CO2	M	S	S	S	M	S	S	M	M	M
CO3	S	S	S	M	S	S	S	M	S	M

CO/PSO	PSO1	PSO2	PSO3	PSO4	PSO5
CO1	3	3	3	3	3
CO2	3	3	3	3	3
Weightage	6	6	6	6	6
Weighted percentage of Course Contribution to POs	3.0	3.0	3.0	3.0	3.0